

TEST NO. 4**CH NO. 4****CHEMISTRY****Minutes****Paper : (Objective type)****12****(Class 9th)-19****Time Allowed:- 15****Maximum Marks:-**

Note:- Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

Q. 1	QUESTIONS	(A)	(B)	(C)	(D)
1.	Which bond present in C ₂ H ₄ molecule?	Single covalent bond	Double covalent bond	Triple covalent bond	Dative covalent bond
2.	The electron deficient molecule is.	NH ₃	BF ₃	N ₂	O ₂
3.	Which one is the non-polar molecule among these?	HCl	HNO ₃	H ₂	H ₂ O
4.	Which compound has hydrogen bonding?	CH ₄	Cl ₂	H ₂	H ₂ O
5.	Which pair has single covalent bond?	CH ₄ and C ₂ H ₄	CH ₄ and CH ₃ Cl	CH ₄ and C ₂ H ₂	CH ₄ and CO ₂
6.	Number of bond pair electrons in nitrogen molecule is.	2	4	6	8
7.	The atom which donates electron pair in Dative covalent bond is called.	Acceptor	Donor	Electronegative	Ionic bonds
8.	Which molecule contains Triple Covalent bond.	H ₂	N ₂	O ₂	Cl ₂
9.	Which molecule is polar	H ₂ O	H ₂	Cl ₂	CO ₂
10.	The density of water at 0° C is:	2.0 g/cm ³	1.00 g/cm ³	0.70 g/cm ³	1.17 g/cm ³
11.	Types of chemical bonds are	1	2	3	4
12.	In Ice Molecule there exist.	Ionic bond	Covalent bond	Co-ordinate bond	Vander wall's process

A	B	C	D	A	B	C	D	A	B	C	D	A	B	C	D	A	B	C	D
1	A	B	C	D	4	A	B	C	D	7	A	B	C	D	10	A	B	C	D
2	A	B	C	D	5	A	B	C	D	8	A	B	C	D	11	A	B	C	D
3	A	B	C	D	6	A	B	C	D	9	A	B	C	D	12	A	B	C	D

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آپشن A, B, C, D کو پین کی سیاہی یا مارکر سے اس طرح پُر کریں کہ سیاہی دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پُر کرنے یا کاٹ کر پُر کرنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

TEST NO.4

CH- 4

CHEMISTRY

(Class 9th)-019

Time Allowed: 1:45

Hours

Paper: I (Essay Type)

Maximum Marks: 48

(PART-I)

2. Write short answers to any FIVE (5) questions:

10

- Difference between Duplet rule and Octet rule?
- Define Ionic bond.
- What is meant by Covalent bond?
- What is meant by bond pair?
- What is meant by Vander Waals forces?
- Define co-ordinate covalent bond.
- What is meant by bonding electrons?
- Why a chemical bond is formed? Explain.

3. Write short answers to any FIVE (5) questions:

10

- What is difference between polar and non- polar covalent bond with examples?
- What is meant by metallic bond?
- Give properties of Ionic compound.
- What is meant by single covalent bond?
- What is meant by triple covalent bond?
- Define double covalent bond?
- Write molecule is polar how?

4. Write short answers to any FIVE (5) questions:

10

- What is meant by Intermolecular forces?
- Why ice floats on the surface of water.
- Give uses of adhesives.
- Write two difference between Ionic and covalent compounds.
- Define dipole-dipole interaction.
- Ionic solids are good conductors of electricity. Why?
- What do you mean by dipole induced forces?
- Why Ionic compound have high melting and boiling points.

(PART - I I)

Note: - Attempt any TWO questions.

5. (a) Define covalent bond. Explain formation of covalent bond in O_2 and N_2 molecules. **5**

(b) Write down the properties of Ionic Compound?

4

6. (a) Define Ionic or electrovalent bond and how sodium chloride and magnesium chloride **5**
are formed?

(b) How a chemical bond is formed? Define the types of covalent bond and give example of **4**
each.

7. (a) Describe the formation of covalent bond. Explain single, double and triple covalent bond **5**
with example.

(b) What are intermolecular forces? Compare these forces with chemical bond forces with **4**
reference to HCl Molecule.



علم کی دُنیا

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