

TEST NO.3**CH- NO. 3****CHEMISTRY****Paper : (Objective type)****Maximum Marks:- 12****(Class 9th)-19****Time Allowed:- 15 Minutes**

Note:- Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

Q.1	QUESTION	(A)	(B)	(C)	(D)
1.	Which one of the following is not a noble gas?	H	He	Ne	Ar
2.	All transition elements are	Metals	Non-metal	Metalloid	Radioactive
3.	The unit of ionization energy is	Nm and pm	KJmol ⁻¹	Newton	Pascal
4.	The no of elements in fourth period of long form of periodic table is.	12	8	18	32
5.	The most electronegative element is	Br	F	Cl	Na
6.	Which is an alkali metal?	Al	Fe	Na	Ca
7.	All the element in a group of periodic table have same _____	Atomic Mass	Mass No.	Valency	Volume
8.	How many blocks are in modern periodic table?	3	4	5	6
9.	Unit of atomic size is	Nm	Pm	KJmol ⁻¹	Both a & b
10.	The unit of ionization energy is.	nm and pm	KJmol ⁻¹	Newton	Pascal
11.	Mendeleev' periodic table was bases upon.	Electronic configuration	Atomic mass	Atomic Number	Mass No.
12.	The number of groups in periodic table is	8	9	18	12

A B C D**A B C D****A B C D****A B C D****A B C D**

1	(A) (B) (C) (D)	4	(A) (B) (C) (D)	7	(A) (B) (C) (D)	10	(A) (B) (C) (D)	13	(A) (B) (C) (D)
2	(A) (B) (C) (D)	5	(A) (B) (C) (D)	8	(A) (B) (C) (D)	11	(A) (B) (C) (D)	14	(A) (B) (C) (D)
3	(A) (B) (C) (D)	6	(A) (B) (C) (D)	9	(A) (B) (C) (D)	12	(A) (B) (C) (D)	15	(A) (B) (C) (D)

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آپشن A,B,C,D کو پین کی سیاہی یا مارکر سے اس طرح پُر کریں کہ سیاہی دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پُر کرنے یا کاٹ کر پُر کرنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

TEST NO.3**CH- 3****CHEMISTRY****Paper: III (Essay Type)****(Class 9th)-019****Time Allowed:1:45 Hours****Maximum Marks:48****(PART – I)****2. Write short answers to any FIVE (5) questions:****10**

- (i) Define electron affinity with an example.
- (ii) Distinguish between period and group in periodic table.
- (iii) Differentiate between lone pair and bond pair of electrons.
- (iv) What is shielding effect? What happens to it in a group?
- (v) What are the transition elements?
- (vi) How many elements are there in 4th period?
- (vii) Why ionization energy of Na is less than Mg?
- (viii) Explain Dobereiner's triads.

3. Write short answers to any FIVE (5) questions:**10**

- (i) What is the Dobereiner's classification of elements?
- (ii) What is the difference between groups and periods in periodic table?
- (iii) Define group number and period number.
- (iv) Define atomic size and atomic radius.
- (v) What is the different Mendeleev's periodic law and Modern periodic law?
- (vi) Define electronegativity. Write electronegativity of Nitrogen, oxygen and Fluorine.
- (vii) Define Mendeleev's periodic law.
- (viii) What are the elements arranged in group 3 to 12 called?

4. Write short answers to any FIVE (5) questions:**10**

- (i) Define lanthanides. In which period it belongs?
- (ii) Why noble gases are not very reactive?
- (iii) Describe Newland of octaves.
- (iv) Write down drawbacks of Mendeleev's periodic table.
- (v) Write two Salient's features of long form of periodic table.
- (vi) Define Ionization Energy.
- (vii) Write any four properties of the long form periodic table.
- (viii) What is meant by periodic function?

Note: - Attempt any TWO questions.

- 5. (a) Define electron affinity. Why it increases in a period and decreases in a group in the periodic table. **5**
- (b) Define atomic size or atomic radius. Write down its unit. Explain its trends in group and period. **4**
- 6. (a) Define Ionization energy. Write down its unit. Explain its trends in periodic table. **5**
- (b) Define Electronegativity. Write down its trends in group and period in the periodic table. **4**
- 7. (a) Discuss any three important features of modern periodic table. **5**
- (b) Define group and explain them in periodic table. **4**