

TEST NO.3**CH- NO. 3****CHEMISTRY****Paper : (Objective type)****Maximum Marks:- 12****(Class 9th)-2020****Time Allowed:- 15 Minutes**

Note:- Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

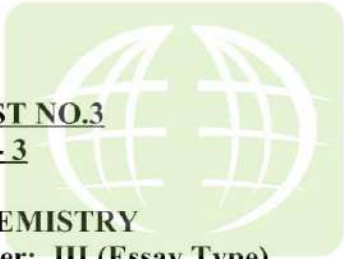
Q.1	QUESTION	(A)	(B)	(C)	(D)
1.	Mendeleev's periodic table was based on.	Electronic configuration	Atomic size	Atomic Number	Mass Number
2.	Modern periodic law was presented by	Dobereiner	Newland	Mendleev e	H. Mosley
3.	Which is not noble gas?	Helium	Hydrogen	Neon	Argon
4.	Transition elements are	All gases	All metals	All non-metals	All metalloids
5.	The base of modern periodic table is.	Mass number	Avogadro's number	Atomic number	Quantum number
6.	Number of elements in normal period are.	18	10	6	32
7.	Which one is shortest period in periodic table?	First period	Second period	Third Peiod	Forth period
8.	Group 17 th belongs.	Halogens	Nobel gases	Alkali metal	None
9.	The distance between the nuclei of two carbon atom is.	154 Pm	140 Pm	110 Pm	115 Pm
10.	Which one of the following halogen has highest electronegativity?	Iodine	Bromine	Chlorine	Fluorine
11.	Electronegativity of carbon is.	2	1	26	4
12.	Which one of the following halogen has lowest electronegativity.	Fluorine	Chlorine	Iodine	Bromine

	A	B	C	D		A	B	C	D		A	B	C	D		A	B	C	D	
1	(A)	(B)	(C)	(D)	4	(A)	(B)	(C)	(D)	7	(A)	(B)	(C)	(D)	10	(A)	(B)	(C)	(D)	13
2	(A)	(B)	(C)	(D)	5	(A)	(B)	(C)	(D)	8	(A)	(B)	(C)	(D)	11	(A)	(B)	(C)	(D)	14
3	(A)	(B)	(C)	(D)	6	(A)	(B)	(C)	(D)	9	(A)	(B)	(C)	(D)	12	(A)	(B)	(C)	(D)	15

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آپشن A, B, C, D کو بیچن کی سی ای یا مار کر سے اس طرح پُر کریں کہ سی ای دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پُر کرنے یا کات کرنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

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TEST NO.3
CH-3



علم كيدنيا
(Class 9th)-2020

Time Allowed:1:45 Hours
Maximum Marks:48

CHEMISTRY
Paper: III (Essay Type)

(PART – I)

- 2. Write short answers to any FIVE (5) questions:** 10
- (i) Explain Dobereiner's Triads.
 - (ii) Define Mendleev's periodic law.
 - (iii) Define Mosely periodic law.
 - (iv) What is the difference between Mendeleev's periodic law and Modern Periodic law?
 - (v) What is meant by periodic function?
 - (vi) Why elements are called S and P block elements?
 - (vii) What is the difference between group and period?
 - (viii) How are elements arranged in 4th period?
- 3. Write short answers to any FIVE (5) questions:** 10
- (i) What is the trend of atomic radius in periods and groups of the Periodic table?
 - (ii) Why size of atom decreases from left to right in the periodic table.
 - (iii) What is trend of ionization energy in period?
 - (iv) Why ionization energy of Na is less than Mg?
 - (v) What is Electronegativity? Writ its unit.
 - (vi) Define Newland octaves law.
 - (vii) Define transition elements.
 - (viii) Write any four properties of the long form periodic table.
- 4. Write short answers to any FIVE (5) questions:** 10
- (i) What are periods? Give an example.
 - (ii) Define Ionization Energy.
 - (iii) Define Shielding effect.
 - (iv) What is difference between short period and long period?
 - (v) What are the elements arranged in group 3 to 12 called?
 - (vi) Why the elements of group 1 and group 2 are called S block elements.
 - (vii) What is the trend of electronegativity in the period?
 - (viii) Give reason for the decrease of electronegativity in down the group.

Note: - Attempt any TWO questions.

- 5.** (a) Define electron affinity. Why it increases in a period and descries in a group in the periodic table. 5
- (b) Define atomic size or atomic radius. Write down its unit. Explain its trends in group and period. 4
- 6.** (a) Define Ionization energy. Write down its unit. Explain its trends in periodic table. 5
- (b) Define Electronegativity. Write down its trends in group and period in the periodic table. 4
- 7.** (a) Discuss any three important features of modern periodic table. 5
- (b) Define group and explain them in periodic table. 4

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