

TEST NO. 7**CH- NO. 6****CHEMISTRY****Maximum Marks:- 12****(Class 9th)-2020****Paper : (Objective type)****Time Allowed:- 15 Minutes**

Note:- Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

Q. No.	Questions	(A)	(B)	(C)	(D)
1.	Brass is the example of which solution in the following.	Liquid in solid	Gas in gas	Liquid in Liquid	Solid in solid
2.	Metal alloy is an examples of.	Liquid in Gas	Gas in liquid	Liquid in solid	Solid in solid
3.	Which thing is a solution of gas in gas?	Smoke	Mist	Fog	Air
4.	Fog is an example of which solution?	Liquid in gas	Gas in Liquid	Solid in gas	Liquid in solid
5.	Example of liquid in liquid is.	Water in milk	Air	Fire	Atom
6.	The example of solution of a solid solute in a solid solvent is.	Fog	Brass	Cheese	Air
7.	If 100 cm ³ of alcohol is dissolved in water and make 100 g solution it is called.	m/m%	m/v%	v/m%	v/v%
8.	Which solution contains more water?	0.5 M	0.25 M	1 M	2 M
9.	Which pair of compounds is soluble?	Ether and water	KCl and water	Benzene and water	Petrol and water
10.	Which will show negligible effect to temperature on its solubility?	KCl	KNO ₃	NaCl	NaNO ₃
11.	Tyndall effect is shown by.	Chalk solution	Jelly	Paints	Sugar solution
12.	Which one is an example of suspension?	Albumin solution	Soap solution	Starch solution	Milk of magnesia

A B C D**A B C D****A B C D****A B C D****A B C D**

1	(A) (B) (C) (D)	4	(A) (B) (C) (D)	7	(A) (B) (C) (D)	10	(A) (B) (C) (D)	13	(A) (B) (C) (D)
2	(A) (B) (C) (D)	5	(A) (B) (C) (D)	8	(A) (B) (C) (D)	11	(A) (B) (C) (D)	14	(A) (B) (C) (D)
3	(A) (B) (C) (D)	6	(A) (B) (C) (D)	9	(A) (B) (C) (D)	12	(A) (B) (C) (D)	15	(A) (B) (C) (D)

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آچشن A, B, C, D کو پین کی سیاہی یا مارکر سے اس طرح پرکھیں کہ سیاہی دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پرکھنے یا کاٹ کر پرکھنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

TEST NO. 7

CH- 6

CHEMISTRY

Paper: VI (Essay Type)

(Class 9th)-2020

Time Allowed: 1:45 Hours

Maximum Marks: 48

(PART-I)

2. Write short answers to any FIVE (5) questions:

10

- (i) Define solution with example.
- (ii) Define aqueous solution. Write its components.
- (iii) What is difference between solute and solvents?
- (iv) Write difference between saturated and unsaturated solution.
- (v) Difference between concentrated solution and dilute solution.
- (vi) Write the name of two non-polar solvents.
- (vii) What is meant by mass/mass% (m/m) %?.
- (viii) What is meant molarity? Also write its formula.

3. Write short answers to any FIVE (5) questions:

10

- (i) How molar solution is prepared?
- (ii) How much amount of KOH required to form 1 molar solution?
- (iii) Define solubility.
- (iv) Write any four examples of colloidal solution.
- (v) Identify as colloids and suspension. Paints, jelly soap solution, milk, milk of magnesia, chalk in water.
- (vi) How will you test whether given solution is colloidal solution or not?
- (vii) What is difference between colloid and suspension?
- (viii) What is Tyndall effect? On what factors it depends.

4. Write short answers to any FIVE (5) questions:

10

- (i) Why we stir paints thoroughly before use?
- (ii) Why do suspension not form the homogeneous mixture?
- (iii) What is meant by "Like dissolves like".
- (iv) What is meant by percentage mass/volume (m/v) %?
- (v) How can you justify that blood is a colloid.
- (vi) Why do the colloids show Tyndall effect?
- (vii) What is meant by true solution?
- (viii) Explain the effect of temperature on solubility.

(PART - II)

Note: - Attempt any TWO questions.

5. (a) Explain how dilute solutions are prepared from concentrated solution. **4**
(b) Define Solubility. Give the general principles of solubility. **5**
6. (a) What are the effects of temperature on solubility? Explain. **4**
(b) Write four characteristics of colloids? **5**
7. (a) When we dissolved 20g of NaCl in 400 cm³ of solution, what will be its molarity? **4**
(b) Comparison of the Characteristics of solution, colloid and suspension. **5**