

TEST – 10
CH- NO. 6 – 8 (2nd half book)
CHEMISTRY
Paper : (Objective type)

(Class 9th)-2020

Time Allowed:- 15 Minutes
Maximum Marks:- 12

Note:- Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

Q.1	Questions	(A)	(B)	(C)	(D)
1.	Metal alloy is an example of.	Liquid in Gas	Gas in Liquid	Liquid in Solid	Solid in Solid
2.	Air is an example of solution.	Solid in solid	Solid in liquid	Gas in Gas	Gas in liquid
3.	If 100 cm ³ of alcohol is dissolved in water and make 100 g solution it is called.	m/m%	m/v%	v/m%	v/v%
4.	Which solution contains more water?	0.5 M	0.25 M	1 M	0.25 M
5.	Tyndall effect is shown by.	Chalk solution	Jelly	Paints	Sugar solution
6.	The oxidation number of chlorine in Potassium chlorate KClO ₃ .	+3	+4	+6	+5
7.	The oxidation number of chromium in K ₂ Cr ₂ O ₇ is.	+2	+6	+14	+7
8.	In which cell spontaneous reaction takes place?	Electrolytic cell	Galvanic Cell	Jaxon Cell	Down's cell
9.	The most common example of corrosion is.	Rusting of iron	Chemical decay	Rusting of aluminum	Rusting of Tin
10.	The example of true solution is.	Starch solution	Tooth paste	Soap solution	Ink in water
11.	Which one is an example of suspension?	Milk of magnesia	Albumin solution	Soap solution	Starch solution
12.	Which is heterogeneous mixture?	Mil	Ink	Milk of magnesia	Sugar solution

A B C D

A B C D

A B C D

A B C D

A B C D

1	(A) (B) (C) (D)	4	(A) (B) (C) (D)	7	(A) (B) (C) (D)	10	(A) (B) (C) (D)	13	(A) (B) (C) (D)
2	(A) (B) (C) (D)	5	(A) (B) (C) (D)	8	(A) (B) (C) (D)	11	(A) (B) (C) (D)	14	(A) (B) (C) (D)
3	(A) (B) (C) (D)	6	(A) (B) (C) (D)	9	(A) (B) (C) (D)	12	(A) (B) (C) (D)	15	(A) (B) (C) (D)

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آہٹن A,B,C,D کو پین کی سیاہی یا مارکر سے اس طرح پڑ کریں کہ سیاہی دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پڑ کرنے یا کاٹ کر پڑ کرنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

TEST NO. 10
CH- 6 – 8 (2nd half book)

CHEMISTRY
Paper: I (Essay Type)

(Class 9th)-2020

Time Allowed: 1:45 Hours

Maximum Marks: 48

2. Write short answers to any Five (5) questions.

10

- (i) Define aqueous solution. Write its components.
- (ii) Difference between solute and solvent?
- (iii) What is meant by mass/mass % (m/m)%?
- (iv) How molar solution is prepared?
- (v) Explain oxidation with example on the basis of electron theory.
- (vi) Define Redox Reaction. Give an example.
- (vii) Identify a strong and weak electrolyte among the following compounds.
CuSO₄, Ca(OH)₂, HCl, HNO₃
- (viii) Difference between strong and weak electrolytes.

3. Write short answers to any Five (5) questions.

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- (i) Difference between electrolytes and non –electrolytes.
- (ii) What are anode and cathode?
- (iii) Which force drives the non-spontaneous reaction to take place?
- (iv) Write down two properties of Galvanic cell.
- (v) What is the function of Salt Bridge?
- (vi) Define Corrosion.
- (viii) Write two uses of sodium.

4. Write short answers to any Five (5) questions.

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- (i) Complete and balance the given chemical reaction. Na + H₂
- (ii) Define metallic character.
- (iii) Define Electropositivity? Explain with example.
- (iv) What do you mean by 24 carat gold?
- (v) Write two uses of Silver.
- (vi) Write the names of noble metals.
- (vii) State the significance of non-metals.
- (viii) Write any two chemical properties of halogens.

(PART - I I)

Note: - Attempt any TWO questions.

5. a) Define solubility. Give the general principles of solubility. **4**
b) Explain oxidation and reduction in terms of loss and gain of electron. **5**
6. a) Compare electrolytic and galvanic cell. **4**
b) Write a note on electrolysis of water. **5**
7. a) What are the effects of temperature on solubility? Explain **4**
b) Write comparison between suspension and colloid. **5**