

**TEST – 11**  
**GRAND TEST**

**CHEMISTRY**

(Class 9<sup>th</sup>)-2020

Time Allowed:- 15 Minutes

Paper : (Objective type)

Maximum Marks:- 12

**Note:-** Write answers to the questions on the objective answer sheet provided. Four possible answers A,B,C and D to each question are given. Which answer you consider correct, fill the circle in front of A,B,C or D with Marker or Pen ink to each question on the answer sheet provided.

Q.1	Questions	(A)	(B)	(C)	(D)
1.	The molecular mass of H <sub>2</sub> SO <sub>4</sub> is.	98 g	98 amu	9.8 g	9.6 amu
2.	All of the following are triatomic molecules except.	H <sub>2</sub>	O <sub>3</sub>	H <sub>2</sub> O	CO <sub>2</sub>
3.	Sub "shell" P" can have maximum number of electrons.	1	4	6	8
4.	The base of modern periodic table is.	Avogadro's number	Atomic number	Quantum number	Mass number
5.	Electronegativity of fluorine is.	4	3.5	2.1	3
6.	Methane is an example of.	Double covalent bond	Dative covalent bond	Single covalent bond	Triple covalent bond
7.	The boiling point of sodium chloride is.	1413 °C	1513 °C	1613 °C	1713 °C
8.	Which instrument is used to measure atmospheric pressure?	Manometer	Barometer	Lactometer	Voltmeter
9.	Example of solution fo solid solute in a solid solvent is.	Fog	Cheese	Air	Brass
10.	Which pair of compounds in soluble?	Ether and water	KCl and water	Benzene and water	Petrol and water
11.	The formula of Rust is.	Fe <sub>2</sub> O <sub>3</sub> .nH <sub>2</sub> O	Fe <sub>2</sub> O <sub>3</sub>	Fe(OH) <sub>3</sub> .nH <sub>2</sub> O	Fe(OH) <sub>3</sub>
12.	The most abundant metal is.	Aluminum	Gold	Platinum	Silver

**A B C D**

**A B C D**

**A B C D**

**A B C D**

**A B C D**

1	(A) (B) (C) (D)	4	(A) (B) (C) (D)	7	(A) (B) (C) (D)	10	(A) (B) (C) (D)	13	(A) (B) (C) (D)
2	(A) (B) (C) (D)	5	(A) (B) (C) (D)	8	(A) (B) (C) (D)	11	(A) (B) (C) (D)	14	(A) (B) (C) (D)
3	(A) (B) (C) (D)	6	(A) (B) (C) (D)	9	(A) (B) (C) (D)	12	(A) (B) (C) (D)	15	(A) (B) (C) (D)

نوٹ: معروضی سوال نامے کو توجہ سے پڑھیں اور ہر MCQ کی درست آجشن A,B,C,D کو پین کی سیاہی یا مارکر سے اس طرح پُر کریں کہ سیاہی دائرے سے باہر نہ نکلے۔ ایک سے زیادہ دائروں کو پُر کرنے یا کاٹ کر پُر کرنے کی صورت میں مذکورہ جواب غلط تصور ہوگا۔

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**TEST NO. 11**

## GRAND TEST

### CHEMISTRY

Paper: I (Essay Type)

(Class 9<sup>th</sup>)-2020

Time Allowed: 1:45 Hours

Maximum Marks: 48

#### (PART-I)

2. Write short answers to any Five (5) questions.

10

- (i) Define Nuclear Chemistry.
- (ii) Difference between matter and mixture.
- (iii) Define homogeneous mixture and heterogeneous mixture.
- (iv) Write two properties of Neutron particles.
- (v) Explain the treatment of cancer by radiotherapy.
- (vi) Define Mendleev's periodic law.
- (vii) What is the trend of atomic radius in periods and groups of the periodic table?
- (viii) What is trend of ionization energy in period?

3. Write short answers to any Five (5) questions.

10

- (i) Define single covalent bond and give one example.
- (ii) Difference between donor atom and acceptor atom.
- (iii) Ice floats on the surface of water. Give reason.
- (iv) What is meant by co-ordinate covalent compounds?
- (v) Define Boiling Point and Melting point.
- (vii) Define allotropy. State allotropes of oxygen.
- (viii) Identify as colloids and suspensions. Paints, Jelly, soap solution, Milk, Milk of magnesia, chalk in water.

4. Write short answers to any Five (5) questions.

10

- (i) Define Redox Reaction. Give an example.
- (ii) Where do the electrons flow from Zn electrode in Daniel's cell?
- (iii) Write the redox reaction taking place during the electroplating of chromium.
- (iv) Define electrochemical cell. Write the names of its types.
- (v) How will you compare the electro positivity of Alkali Metals and Alkaline Earth metals?
- (vi) Write down two uses of Gold.
- (vii) Write the names of noble metals.
- (viii) Describe the non-metallic character in groups and period of periodic table.

#### (PART - II)

Note: - Attempt any TWO questions.

5.
  - a) Explain types of molecules in detail and give examples. 4
  - b) Give postulates of Bohr's theory. 5
6.
  - a) What are covalent compounds? Describe properties of covalent compounds. 4
  - b) Define Boyle's law and verify it with an experiment 5
7.
  - a) Write the four characteristics of colloids. 4
  - b) Describe the rules for assigning the oxidation number 5